



## Balancing Chemical Equations 2

### EXERCISES

Balance the following equations by writing in coefficients in front of the molecules.

- 1) \_\_\_ Mg + \_\_\_ N<sub>2</sub> → \_\_\_ Mg<sub>3</sub>N<sub>2</sub>
- 2) \_\_\_ Ca(ClO<sub>3</sub>)<sub>2</sub> → \_\_\_ CaCl<sub>2</sub> + \_\_\_ O<sub>2</sub>
- 3) \_\_\_ K + \_\_\_ H<sub>2</sub>O → \_\_\_ KOH + \_\_\_ H<sub>2</sub>
- 4) \_\_\_ HCl + \_\_\_ Ba(OH)<sub>2</sub> → \_\_\_ BaCl<sub>2</sub> + \_\_\_ H<sub>2</sub>O
- 5) \_\_\_ Pb(NO<sub>3</sub>)<sub>2</sub> + \_\_\_ NaCl → \_\_\_ PbCl<sub>2</sub> + \_\_\_ NaNO<sub>3</sub>
- 6) \_\_\_ As<sub>2</sub>S<sub>3</sub> + \_\_\_ HCl → \_\_\_ AsCl<sub>3</sub> + \_\_\_ H<sub>2</sub>S
- 7) \_\_\_ SO<sub>2</sub> + \_\_\_ O<sub>2</sub> → \_\_\_ SO<sub>3</sub>
- 8) \_\_\_ NH<sub>4</sub>NO<sub>3</sub> → \_\_\_ N<sub>2</sub>O + \_\_\_ H<sub>2</sub>O
- 9) \_\_\_ Al + \_\_\_ H<sub>2</sub>SO<sub>4</sub> → \_\_\_ Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> + \_\_\_ H<sub>2</sub>
- 10) \_\_\_ Zn + \_\_\_ HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> → \_\_\_ Zn(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub> + \_\_\_ H<sub>2</sub>
- 11) \_\_\_ C<sub>5</sub>H<sub>12</sub> + \_\_\_ O<sub>2</sub> → \_\_\_ CO<sub>2</sub> + \_\_\_ H<sub>2</sub>O
- 12) \_\_\_ C<sub>6</sub>H<sub>14</sub> + \_\_\_ O<sub>2</sub> → \_\_\_ CO<sub>2</sub> + \_\_\_ H<sub>2</sub>O
- 13) \_\_\_ BaCl<sub>2</sub> + \_\_\_ (NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub> → \_\_\_ BaCO<sub>3</sub> + \_\_\_ NH<sub>4</sub>Cl
- 14) \_\_\_ Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> + \_\_\_ H<sub>3</sub>PO<sub>4</sub> → \_\_\_ Ca(H<sub>2</sub>PO<sub>4</sub>)<sub>2</sub>
- 15) \_\_\_ Mg + \_\_\_ B<sub>2</sub>O<sub>3</sub> → \_\_\_ MgO + \_\_\_ B
- 16) \_\_\_ Al + \_\_\_ MnO<sub>2</sub> → \_\_\_ Mn + \_\_\_ Al<sub>2</sub>O<sub>3</sub>
- 17) \_\_\_ Al<sub>2</sub>S<sub>3</sub> + \_\_\_ H<sub>2</sub>O → \_\_\_ Al(OH)<sub>3</sub> + \_\_\_ H<sub>2</sub>S
- 18) \_\_\_ C<sub>5</sub>H<sub>11</sub>OH + \_\_\_ O<sub>2</sub> → \_\_\_ CO<sub>2</sub> + \_\_\_ H<sub>2</sub>O
- 19) \_\_\_ H<sub>2</sub>O<sub>2</sub> → \_\_\_ H<sub>2</sub>O + \_\_\_ O<sub>2</sub>
- 20) \_\_\_ Ca + \_\_\_ H<sub>2</sub>O → \_\_\_ Ca(OH)<sub>2</sub> + \_\_\_ H<sub>2</sub>



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## SOLUTIONS

